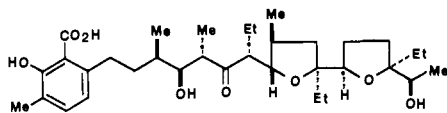


- (14) K. B. Sharpless and R. C. Michaelson, *J. Am. Chem. Soc.*, **95**, 6136 (1973); S. Tanaka, H. Yamamoto, H. Nozaki, K. B. Sharpless, R. C. Michaelson, and J. D. Cutting, *J. Am. Chem. Soc.*, **96**, 5254 (1974).
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- (16) Numbering in this paper corresponds to that of lasalocids: J. W. Westley, *J. Antibiot.*, **29**, 584 (1976).
- (17) This acetylation was necessary to avoid formation of the undesired tetrahydrofuran ring at this stage.
- (18) Similar conditions were used to open the epoxide ring of Juvenile hormone: K. Nakanishi, D. A. Schooley, M. Koreeda, J. Dillon, *Chem. Commun.*, 1235 (1971).
- (19) D. H. R. Barton, P. D. Magnus, G. Smith, G. Streckert, and D. Zurr, *J. Chem. Soc., Perkin Trans. 1*, 542 (1972).
- (20) We are indebted to Dr. J. W. Westley for samples of lasalocid A (1) and isolasalocid A (i).



i, isolasalocid A

- (21) This substance was prepared by following the synthetic method of 2-acetoxy-2-methyl-6-carboethoxy-3,5-cyclohexadien-1-one: F. Wessely, E. Zbrlral, and H. Sturm, *Chem. Ber.*, **93**, 2840 (1960).
- (22) This optically active bromide was synthesized as follows: (1) $\text{CH}_3\text{CH}=\text{CHCH}_2\text{OH}/\text{CH}_3(\text{OEt})_3$, 160 °C; (2) $\text{NaOH}/\text{aqueous dioxane}$, room temper-

- ature; (3) optical resolution as the α -methylbenzylamine salt; (4) $\text{LiAlH}_4/\text{Et}_2\text{O}$, room temperature; (5) $\text{MsCl}/\text{Et}_3\text{N}/\text{CH}_2\text{Cl}_2$, 0 °C; (6) LiBr/DMF , 100 °C.
- (23) The authentic aldehyde **14** was prepared from natural lasalocid A in two steps: (1) $\text{C}_6\text{H}_5\text{CH}_2\text{Br}/\text{K}_2\text{CO}_3/\text{dioxane}$, 80 °C; (2) 230 °C (0.02 mmHg).³
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- (25) The Aldol condensation between **12** and dibenzylated aldehyde (cf. **14**) was as good as the one reported.
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T. Nakata, G. Schmid, B. Vranesic, M. Okigawa
T. Smith-Palmer, Y. Kishi*

Department of Chemistry, Harvard University
Cambridge, Massachusetts 02138

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Additions and Corrections

Electrochemical Synthesis and Structural Characterization of the Iron-Sulfur Cluster Cation $[(\eta\text{-C}_5\text{H}_5)_2\text{Fe}_2(\text{S}_2)(\text{SC}_2\text{H}_5)_2]^+$ [*J. Am. Chem. Soc.*, **98**, 1980 (1976)]. By P. J. VERGAMINI,* R. R. RYAN, and G. J. KUBAS, The University of California, Los Alamos Scientific Laboratory, Los Alamos, New Mexico 87545.

An error has been found in the atomic position parameter table published as Supplementary Material. A revised table is deposited herewith.

Supplementary Material Available: A listing of atomic positional parameters (Table I) (1 page). Ordering information is given on any current masthead page.

Ionization of Group 6 and 7 Protonic Acids in Dimethyl Sulfide [*J. Am. Chem. Soc.*, **99**, 808 (1977)]. By EDWARD M. ARNETT* and LEONARD E. SMALL, Department of Chemistry, University of Pittsburgh, Pittsburgh, Pennsylvania 15260.

Page 814: The captions for Figures 3 and 4 are reversed.

Add as ref 7a: F. A. Long and P. Ballinger, "Electrolytes", B. Pesce, Ed., Pergamon Press, New York, N.Y., 1962, pp. 152-165.

An Anionic Equivalent of the Friedel-Crafts Cycloacylation [*J. Am. Chem. Soc.*, **99**, 4822 (1977)]. By R. J. BOATMAN, B. J. WHITLOCK, and H. W. WHITLOCK, JR.,* Department of Chemistry, University of Wisconsin—Madison, Madison, Wisconsin 53706.

It has been called to our attention that our references 4 and 6 to Professor W. E. Parham's work in this area were not very explicit. Since our primary concern is with the synthesis of anthracyclonones, we did not give a detailed literature review,

but we would point out now that a comparable synthesis of 1-indanone by him is described in the paper referred to in ref 6, which also indicates the potential generality and utility of this synthetic approach.

Magnesium Bromide-Tetrahydrofuran Complexes: $\text{MgBr}_2(\text{C}_4\text{H}_8\text{O})_2$, $\text{MgBr}_2(\text{C}_4\text{H}_8\text{O})_3$, $\text{MgBr}_2(\text{C}_4\text{H}_8\text{O})_4$, and $\text{MgBr}_2(\text{C}_4\text{H}_8\text{O})_4(\text{H}_2\text{O})_2$. A Reagent for the Preparation of Anhydrous Magnesium Phosphodiester Salts [*J. Am. Chem. Soc.*, **99**, 5285 (1977)]. By FAUSTO RAMIREZ,* RAGHUPATHY SARMA,* YU FEN CHAW, TERENCE M. MCCAFFREY, JAMES F. MARECEK, BRIAN MCKEEVER, and DAVID NIERMAN, Chemistry and Biochemistry Departments, State University of New York at Stony Brook, Stony Brook, New York 11794.

Experimental Section, page 5287, line 11: For "The solution was concentrated to ca. 200 mL . . ." read "The solution was concentrated to ca. 50 mL . . .".

Orientation in Nucleophilic Substitution at the Cycloheptatrienone Nucleus: Failure of Predictions from Either Electron Spin Resonance Data or Molecular Orbital Treatments [*J. Am. Chem. Soc.*, **99**, 5997 (1977)]. By MARINO CAVAZZA, M. PERLA COLOMBINI, MASSIMO MARTINELLI, LAMBERTO NUCCI, LUCIO PARDI, FRANCESCO PIETRA,* and SERGIO SANTUCCI, Department of Chemistry and GNSM, Istituto di Fisica, Università di Pisa, 56100 Pisa, Italy, and Facoltà di Scienze, Libera Università di Trento, 38050 Povo (Trento), Italy.

UV data for 3-thiomethyltropone are correct as they stand in Table VI while those reported in the Experimental Section refer instead to 2-thiomethyltropone.